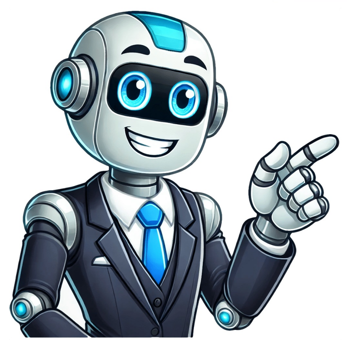

Electron configuration for selenium

Selenium Electron Configuration: Chemical element selenium has the symbol Se and its atomic number is 34. It comes in the category of non-metals. And its properties are lies between the above and below elements of selenium in the periodic table i.e. Sulphur and tellurium. Properties of selenium are also similar to Arsenic. Selenium Electron Configuration The electronic configuration is defined as the distribution of a number of electrons in the atom or molecule. Electron Configuration For Selenium Selenium consists of 34 electrons distribution in its 4 orbits. So electronic configuration of selenium define as: 1s22s22p63s23p63d104s2 3d104p4. Or The electronic configuration can also be represented with the help of a full electron distribution element that is Argon/ Ar. In this case it is written as [Ar] 3d104s24p4 What is the Electron Configuration of Selenium? The electronic configuration of selenium is represented as 1s22s22p63s23p63d104s24p4 or [Ar] 3d104s24p4 Selenium Valence Electrons Valence electrons are the electrons which are located in the outermost shell of the atom or molecule. Selenium has 6 electrons in its outermost shell i.e. 2 electrons in s orbit and 4 electrons in the p orbit. How Many Valence Electrons Does Selenium Have Selenium has 6 electrons located in the outermost shell. The number of valence electrons is defined as the number of electrons present in the outermost shell of the atom. So, in the case of selenium, there are 6 valence electrons. Selenium is the 34th element in the periodic table and the symbol is 'Se'. Selenium has an atomic number of 34, which means that its atom has 34 electrons around its nucleus.The electron configuration of selenium is 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p4 which means that the first two electrons enter the 1s orbital. Since the 1s orbital can hold only two electrons the next two will enter the 2s orbital. The next six electrons enter the 2p subshell. The p subshell can hold a maximum of six electrons. So first we put six electrons in the 2p subshell and then the next two electrons in the 3s orbital.Since the 3s is now full, the electrons will move to the 3p subshell, where the next six electrons will enter. The 3p subshell is now full. Consequently, the following two electrons will enter the 4s orbital. Since the 4s orbital is full, the next ten electrons will move into the 3d subshell. The 3d subshell is now full. Consequently, the remaining four electrons will enter the 4p subshell. Hence, the electron configuration of selenium will be 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p4.The electron configuration of selenium refers to the arrangement of electrons in the selenium atom's orbitals. It describes how electrons are distributed among the various atomic orbitals and energy levels, and provides a detailed map of where each electron is likely to be found.To understand the mechanism of selenium electron configuration, you need to understand two basic things. These are orbits and orbitals. Also, you can arrange electrons in those two ways. In this article, I have discussed all the necessary points to understand the mechanism of selenium electron configuration, including its ground state, noble gas notation, orbital diagram with Bohr model, valence electrons, and configurations for Se^{2-}. I hope this will be helpful in your study.Selenium(Se) electron configuration (Bohr model)Scientist Niels Bohr was the first to give an idea of the atom's orbit. He provided a model of the atom in 1913 and provided a complete idea of orbit in that model.The electrons of the atom revolve around the nucleus in a certain circular path. These circular paths are called orbits (shells or energy levels). These orbits are expressed by n. [n = 1,2,3,4 . . . The serial number of the orbit]The name of the first orbit is K, L is the second, M is the third, and N is the name of the fourth orbit. The electron holding capacity of each orbit is 2n2.Shell Number (n)Shell NameElectrons Holding Capacity (2n2)1K22L83M184N32Electron holding capacity of shellsLet, n = 1 for K orbit. So, the maximum electron holding capacity in the K orbit is 2n2 = 2 × 12 = 2 electrons.n = 2, for L orbit. The maximum electron holding capacity in the L orbit is 2n2 = 2 × 22 = 8 electrons.n=3 for M orbit. The maximum electron holding capacity in the M orbit is 2n2 = 2 × 32 = 18 electrons.n=4 for N orbit. The maximum electron holding capacity in N orbit is 2n2 = 2 × 42 = 32 electrons.Therefore, the maximum electron holding capacity in the first shell is two, the second shell is eight and the 3rd shell can have a maximum of eighteen electrons.Position of selenium(Se) in the periodic tableThe atomic number is the number of electrons in that element. The atomic number of selenium is 34. That is, the number of electrons in selenium is thirty-four. Therefore, the selenium atom will have two electrons in the first shell, eight in the 2nd orbit, eighteen electrons in the 3rd shell, and the remaining six electrons in the fourth shell. Therefore, the order of the number of electrons in each shell of the selenium(Se) atom is 2, 8, 18, 6.The Bohr atomic model has many limitations. In the Bohr atomic model, the electrons can only be arranged in different shells but the exact position, orbital shape, and spin of the electron cannot be determined.Also, electrons can be arranged correctly from 1 to 18 elements. The electron arrangement of any element with atomic number greater than 18 cannot be accurately determined by the Bohr atomic model following the 2n2 formula. We can overcome all limitations of the Bohr model following the electron configuration through orbital.Atomic energy shells are subdivided into sub-energy levels. These sub-energy levels are also called orbital. The most probable region of electron rotation around the nucleus is called the orbital.The sub-energy levels depend on the azimuthal quantum number. It is expressed by 'l'. The value of 'l' is from 0 to (n – 1). The sub-energy levels are known as s, p, d, and f.Orbit NumberValue of 'l'Number of subshellsNumber of orbitalsSubshell nameElectrons holding capacityElectron configuration10111s21s22012132s2p262s2 2p6301231353s3p3d26103s2 3p6 3d1040123413574s4p4d4f2610144s2 4p6 4d10 4f14Orbital number of the subshellIf n = 1,(n – 1) = (1–1) = 0Therefore, the value of 'l' is 0. So, the sub-energy level is 1s.If n = 2,(n – 1) = (2–1) = 1.Therefore, the value of 'l' is 0, 1. So, the sub-energy levels are 2s, and 2p.If n = 3,(n – 1) = (3–1) = 2.Therefore, the value of 'l' is 0, 1, 2. So, the sub-energy levels are 3s, 3p, and 3d.If n = 4,(n – 1) = (4–1) = 3Therefore, the value of 'l' is 0, 1, 2, 3. So, the sub-energy levels are 4s, 4p, 4d, and 4f.If n = 5,(n – 1) = (n – 5) = 4.Therefore, l = 0,1,2,3,4. The number of sub-shells will be 5 but 4s, 4p, 4d, and 4f in these four subshells it is possible to arrange the electrons of all the elements of the periodic table.Sub-shell nameName sourceValue of 'l'Value of 'm'(0 to ± l)Number of orbital (2l+1)Electrons holding capacity2(2l+1)sSharp0012pPrincipal1−1, 0, +136dDiffuse2−2, −1, 0, +1, +2510fFundamental3−3, −2, −1, 0, +1, +2, +3714Number of electrons in the orbitalThe orbital number of the s-subshell is one, three in the p-subshell, five in the d-subshell, and seven in the f-subshell. Each orbital can have a maximum of two electrons.The sub-energy level 's' can hold a maximum of two electrons, 'p' can hold a maximum of six electrons, 'd' can hold a maximum of ten electrons, and 'f' can hold a maximum of fourteen electrons.Aufbau is a German word, which means building up. The main proponents of this principle are scientists Niels Bohr and Pauli. The Aufbau method is to do electron configuration through the sub-energy level.The Aufbau principle is that the electrons present in the atom will first complete the lowest energy orbital and then gradually continue to complete the higher energy orbital.Electron Configuration Mechanism Through Aufbau PrincipalThe energy of an orbital is calculated from the value of the principal quantum number 'n' and the azimuthal quantum number 'l'. The orbital for which the value of (n + l) is lower is the low energy orbital and the electron will enter that orbital first.OrbitalOrbit (n)Azimuthal quantum number (l)Orbital energy (n + l)3d3254s404Energy of orbitalHere, the energy of 4s orbital is less than that of 3d. So, the electron will enter the 4s orbital first and enter the 3d orbital when the 4s orbital is full. Following the Aufbau principle, the sequence of entry of electrons into orbitals is 1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p.Therefore, the complete electron configuration for selenium should be written as 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p4.Selenium electron configurationNote: The unabbreviated electron configuration of selenium is [Ar] 3d10 4s2 4p4. When writing an electron configuration, you have to write serially.Atoms can jump from one orbital to another orbital in an excited state. This is called a quantum jump. The ground state electron configuration of selenium is 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4p4. In the selenium ground-state electron configuration, the last four electrons of the 4p orbital are located in the 4px(2), 4py and 4pz orbitals.The p-subshell has three orbitals. The orbitals are px, py, and pz and each orbital can have a maximum of two electrons. Then the correct electron configuration of selenium in the ground state will be 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4px2 4py1 4pz1. This electron configuration shows that the last shell of the selenium atom has two unpaired electrons. So in this case, the valency of selenium is 2.Selenium excited state electron configuration and orbital diagramWhen the selenium atom is excited, then the selenium atom absorbs energy. As a result, an electron in the 4px orbital jumps to the 4dxy1 orbital. We already know that the d-subshell has five orbitals.The orbitals are dxy, dyz, dzx, dx2-y2 and dz2 and each orbital can have a maximum of two electrons. Therefore, the electron configuration of selenium(Se*) in an excited state will be 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4px1 4py1 4pz1 4dxy1. The valency of the element is determined by electron configuration in the excited state. Here, selenium has four unpaired electrons. So, the valency of selenium is 4.The electron configuration of selenium shows that the last shell of selenium has six electrons. Therefore, the valence electrons of selenium are six. The elements that have 5, 6, or 7 electrons in the last shell receive the electrons in the last shell during bond formation. The elements that receive electrons and form bonds are called anions.Atomic number of selenium, atomic weight and chargeDuring the formation of a bond, the last shell of selenium receives two electrons and turns into a selenium ion(Se2-). That is, selenium is an anion element.Se + 2e– → Se2-The electron configuration of selenium ion(Se2-) is 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4p6. This electron configuration shows that the selenium ion(Se2-) acquired the electron configuration of krypton. Selenium atoms exhibit -2, +2, +4, +6 oxidation states. The oxidation state of the element changes depending on the bond formation. Farhan SadikHi, I'm Farhan Sadik. I've always been captivated by chemistry since my school days and pursued extensive research during college, especially on the periodic table. As a full-time chemistry writer on Valenceelectrons.com, my mission is to share the knowledge I've gained about electron configuration, valence electrons, and atomic properties. I believe that quality education should be accessible to all, and I hope to empower learners worldwide to explore the wonders of chemistry. Germanium[Ar] 3d10 4s2 4p2Selenium[Ar] 3d10 4s2 4p4Antimony[Kr] 4d10 5s2 5p3Tellurium[Kr] 4d10 5s2 5p4Tantalum[Xe] 4f14 5d3 6s2Tungsten[Xe] 4f14 5d4 6s2Platinum[Xe] 4f14 5d9 6s1Mercury[Xe] 4f14 5d10 6s2Rutherfordium[Rn] 5f14 6d2 7s2Seaborgium[Rn] 5f14 6d4 7s2Meitnerium[Rn] 5f14 6d7 7s2Darmstadtium[Rn] 5f14 6d8 7s2Roentgenium[Rn] 5f14 6d9 7s2Copernicium[Rn] 5f14 6d10 7s2Nihonium[Rn] 5f14 6d10 7s2 7p1Flerovium[Rn] 5f14 6d10 7s2 7p2Moscovium[Rn] 5f14 6d10 7s2 7p3Livermorium[Rn] 5f14 6d10 7s2 7p4Tennessine[Rn] 5f14 6d10 7s2 7p5Oganesson[Rn] 5f14 6d10 7s2 7p6Gadolinium[Xe] 4f7 5d1 6s2Lutetium[Xe] 4f14 5d1 6s2Protactinium[Rn] 5f2 6d1 7s2Neptunium[Rn] 5f4 6d1 7s2Lawrencium[Rn] 5f14 7s2 7p1 Selenium is the 34th element in the periodic table and the symbol is 'Se'. Selenium has an atomic number of 34, which means that its atom has 34 electrons around its nucleus.The electron configuration of selenium is 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p4 which means that the first two electrons enter the 1s orbital. Since the 1s orbital can hold only two electrons the next two will enter the 2s orbital. The next six electrons enter the 2p subshell. The p subshell can hold a maximum of six electrons. So first we put six electrons in the 2p subshell and then the next two electrons in the 3s orbital.Since the 3s is now full, the electrons will move to the 3p subshell, where the next six electrons will enter. The 3p subshell is now full. Consequently, the following two electrons will enter the 4s orbital. Since the 4s orbital is full, the next ten electrons will move into the 3d subshell. The 3d subshell is now full. Consequently, the remaining four electrons will enter the 4p subshell. Hence, the electron configuration of selenium will be 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p4.The electron configuration of selenium refers to the arrangement of electrons in the selenium atom's orbitals. It describes how electrons are distributed among the various atomic orbitals and energy levels, and provides a detailed map of where each electron is likely to be found.To understand the mechanism of selenium electron configuration, you need to understand two basic things. These are orbits and orbitals. Also, you can arrange electrons in those two ways. In this article, I have discussed all the necessary points to understand the mechanism of selenium electron configuration, including its ground state, noble gas notation, orbital diagram with Bohr model, valence electrons, and configurations for Se^{2-}. I hope this will be helpful in your study.Selenium(Se) electron configuration (Bohr model)Scientist Niels Bohr was the first to give an idea of the atom's orbit. He provided a model of the atom in 1913 and provided a complete idea of orbit in that model.The electrons of the atom revolve around the nucleus in a certain circular path. These circular paths are called orbits (shells or energy levels). These orbits are expressed by n. [n = 1,2,3,4 . . . The serial number of the orbit]The name of the first orbit is K, L is the second, M is the third, and N is the name of the fourth orbit. The electron holding capacity of each orbit is 2n2.Shell Number (n)Shell NameElectrons Holding Capacity (2n2)1K22L83M184N32Electron holding capacity of shellsLet, n = 1 for K orbit. So, the maximum electron holding capacity in the K orbit is 2n2 = 2 × 12 = 2 electrons.n = 2, for L orbit. The maximum electron holding capacity in the L orbit is 2n2 = 2 × 22 = 8 electrons.n=3 for M orbit. The maximum electron holding capacity in the M orbit is 2n2 = 2 × 32 = 18 electrons.n=4 for N orbit. The maximum electron holding capacity in N orbit is 2n2 = 2 × 42 = 32 electrons.Therefore, the maximum electron holding capacity in the first shell is two, the second shell is eight and the 3rd shell can have a maximum of eighteen electrons.Position of selenium(Se) in the periodic tableThe atomic number is the number of electrons in that element. The atomic number of selenium is 34. That is, the number of electrons in selenium is thirty-four. Therefore, the selenium atom will have two electrons in the first shell, eight in the 2nd orbit, eighteen electrons in the 3rd shell, and the remaining six electrons in the fourth shell. Therefore, the order of the number of electrons in each shell of the selenium(Se) atom is 2, 8, 18, 6.The Bohr atomic model has many limitations. In the Bohr atomic model, the electrons can only be arranged in different shells but the exact position, orbital shape, and spin of the electron cannot be determined.Also, electrons can be arranged correctly from 1 to 18 elements. The electron arrangement of any element with atomic number greater than 18 cannot be accurately determined by the Bohr atomic model following the 2n2 formula. We can overcome all limitations of the Bohr model following the electron configuration through orbital.Atomic energy shells are subdivided into sub-energy levels. These sub-energy levels are also called orbital. The most probable region of electron rotation around the nucleus is called the orbital.The sub-energy levels depend on the azimuthal quantum number. It is expressed by 'l'. The value of 'l' is from 0 to (n – 1). The sub-energy levels are known as s, p, d, and f.Orbit NumberValue of 'l'Number of subshellsNumber of orbitalsSubshell nameElectrons holding capacityElectron configuration10111s21s22012132s2p262s2 2p6301231353s3p3d26103s2 3p6 3d1040123413574s4p4d4f2610144s2 4p6 4d10 4f14Orbital number of the subshellIf n = 1,(n – 1) = (1–1) = 0Therefore, the value of 'l' is 0. So, the sub-energy level is 1s.If n = 2,(n – 1) = (2–1) = 1.Therefore, the value of 'l' is 0, 1. So, the sub-energy levels are 2s, and 2p.If n = 3,(n – 1) = (3–1) = 2.Therefore, the value of 'l' is 0, 1, 2. So, the sub-energy levels are 3s, 3p, and 3d.If n = 4,(n – 1) = (4–1) = 3Therefore, the value of 'l' is 0, 1, 2, 3. So, the sub-energy levels are 4s, 4p, 4d, and 4f.If n = 5,(n – 1) = (n – 5) = 4.Therefore, l = 0,1,2,3,4. The number of sub-shells will be 5 but 4s, 4p, 4d, and 4f in these four subshells it is possible to arrange the electrons of all the elements of the periodic table.Sub-shell nameName sourceValue of 'l'Value of 'm'(0 to ± l)Number of orbital (2l+1)Electrons holding capacity2(2l+1)sSharp0012pPrincipal1−1, 0, +136dDiffuse2−2, −1, 0, +1, +2510fFundamental3−3, −2, −1, 0, +1, +2, +3714Number of electrons in the orbitalThe orbital number of the s-subshell is one, three in the p-subshell, five in the d-subshell, and seven in the f-subshell. Each orbital can have a maximum of two electrons.The sub-energy level 's' can hold a maximum of two electrons, 'p' can hold a maximum of six electrons, 'd' can hold a maximum of ten electrons, and 'f' can hold a maximum of fourteen electrons.Aufbau is a German word, which means building up. The main proponents of this principle are scientists Niels Bohr and Pauli. The Aufbau method is to do electron configuration through the sub-energy level.The Aufbau principle is that the electrons present in the atom will first complete the lowest energy orbital and then gradually continue to complete the higher energy orbital.Electron Configuration Mechanism Through Aufbau PrincipalThe energy of an orbital is calculated from the value of the principal quantum number 'n' and the azimuthal quantum number 'l'. The orbital for which the value of (n + l) is lower is the low energy orbital and the electron will enter that orbital first.OrbitalOrbit (n)Azimuthal quantum number (l)Orbital energy (n + l)3d3254s404Energy of orbitalHere, the energy of 4s orbital is less than that of 3d. So, the electron will enter the 4s orbital first and enter the 3d orbital when the 4s orbital is full. Following the Aufbau principle, the sequence of entry of electrons into orbitals is 1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p.Therefore, the complete electron configuration for selenium should be written as 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p4.Selenium electron configurationNote: The unabbreviated electron configuration of selenium is [Ar] 3d10 4s2 4p4. When writing an electron configuration, you have to write serially.Atoms can jump from one orbital to another orbital in an excited state. This is called a quantum jump. The ground state electron configuration of selenium is 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4p4. In the selenium ground-state electron configuration, the last four electrons of the 4p orbital are located in the 4px(2), 4py and 4pz orbitals.The p-subshell has three orbitals. The orbitals are px, py, and pz and each orbital can have a maximum of two electrons. Then the correct electron configuration of selenium in the ground state will be 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4px2 4py1 4pz1. This electron configuration shows that the last shell of the selenium atom has two unpaired electrons. So in this case, the valency of selenium is 2.Selenium excited state electron configuration and orbital diagramWhen the selenium atom is excited, then the selenium atom absorbs energy. As a result, an electron in the 4px orbital jumps to the 4dxy1 orbital. We already know that the d-subshell has five orbitals.The orbitals are dxy, dyz, dzx, dx2-y2 and dz2 and each orbital can have a maximum of two electrons. Therefore, the electron configuration of selenium(Se*) in an excited state will be 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4px1 4py1 4pz1 4dxy1. The valency of the element is determined by electron configuration in the excited state. Here, selenium has four unpaired electrons. So, the valency of selenium is 4.The electron configuration of selenium shows that the last shell of selenium has six electrons. Therefore, the valence electrons of selenium are six. The elements that have 5, 6, or 7 electrons in the last shell receive the electrons in the last shell during bond formation. The elements that receive electrons and form bonds are called anions.Atomic number of selenium, atomic weight and chargeDuring the formation of a bond, the last shell of selenium receives two electrons and turns into a selenium ion(Se2-). That is, selenium is an anion element.Se + 2e– → Se2-The electron configuration of selenium ion(Se2-) is 1s2 2s2 2p6 3s2 3p6 3d10 4s2 4p6. This electron configuration shows that the selenium ion(Se2-) acquired the electron configuration of krypton. Selenium atoms exhibit -2, +2, +4, +6 oxidation states. The oxidation state of the element changes depending on the bond formation. Farhan SadikHi, I'm Farhan Sadik. I've always been captivated by chemistry since my school days and pursued extensive research during college, especially on the periodic table. As a full-time chemistry writer on Valenceelectrons.com, my mission is to share the knowledge I've gained about electron configuration, valence electrons, and atomic properties. I believe that quality education should be accessible to all, and I hope to empower learners worldwide to explore the wonders of chemistry. Selenium is a chemical element of the periodic table with chemical symbol Se and atomic number 34 with an atomic weight of 78.9718 u and is classed as nonmetal and is part of group 16 (oxygen group). Selenium is solid at room temperature. Arsenic Periodic table Bromine SymbolSeAtomic number34Group16 (Oxygen group)Period4BlockpClassificationNonmetal AppearanceBlack, red, and gray (not pictured) allotropesColor Gray Number of protons34 p+Number of neutrons45 n0Number of electrons34 e- From Wikipedia, the free encyclopediaSelenium is a chemical element with symbol Se and atomic number 34. It is a nonmetal with properties that are intermediate between those of its periodic table column-adjacent chalcogen elements sulfur and tellurium. It rarely occurs in its elemental state in nature, or as pure ore compounds. Phase at STPSolidDensity4.81 g/cm3Atomic weight78.9718 u Melting point494 K220.85 °C429.53 °FBoiling point958 K684.85 °C1264.73 °FHeat of vaporization26.32 kJ/mol Electronegativity (Pauling Scale)2.55Electron affinity194.959 kJ/molOxidation states−2, −1, +1, +2, +3, +4, +5, +6(a strongly acidic oxide)Ionization energies941 kJ/mol 2045 kJ/mol 2973.7 kJ/mol 4144 kJ/mol 6590 kJ/mol 7880 kJ/mol 14990 kJ/mol Electron configuration Shorthand configuration[Ar] 3d10 4s2 4p4 Electron configuration Full configuration1s2 2s2 2p6 3s2 3p6 3d10 4s2 4p4 Electron configuration chartElectrons per shell2, 8, 18, 6Valence electrons 6Valency electrons 2,4,6Bohr modelElectron shell for Selenium, created by Injosoft ABSeFigure: Shell diagram of Selenium (Se) atom.Orbital Diagram Discovery and first isolationJöns Jakob Berzelius, Johan Gottlieb Gahn (1817)Discovery of seleniumSelenium (Greek σελήνη selene meaning "Moon") was discovered in 1817 by Jöns Jacob Berzelius and Johan Gottlieb Gahn. Both chemists owned a chemistry plant near Gripsholm, Sweden, producing sulfuric acid by the lead chamber process. The pyrite from the Falun Mine created a red precipitate in the lead chambers which was presumed to be an arsenic compound, so the pyrite's use to make acid was discontinued. Berzelius and Gahn wanted to use the pyrite and they also observed that the red precipitate gave off a smell like horseradish when burned. This smell was not typical of arsenic, but a similar odor was known from tellurium compounds. Hence, Berzelius's first letter to Alexander Marcet stated that this was a tellurium compound. However, the lack of tellurium compounds in the Falun Mine minerals eventually led Berzelius to reanalyze the red precipitate, and in 1818 he wrote a second letter to Marcet describing a newly found element similar to sulfur and tellurium. Because of its similarity to tellurium, named for the Earth, Berzelius named the new element after the Moon. CAS Number7782-49-2ChemSpider ID4885617EC number231-957-4PubChem CID Number6326970